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This type of analysis is known as qualitative analysis. An anion is a negatively charged ion. 2. Why is deionised water used in all of the tests? Since water is a good solvent, tap water contains many dissolved ions. These ions could interfere with the tests. Deionised water is used because it contains no dissolved ions. 3. What is a precipitate? In the test for the chloride ion, how can the presence of silver chloride be confirmed? A precipitate is a material that settles out of solution. In the test for the chloride ion, the presence of silver chloride is confirmed by adding dilute ammonia solution to the silver chloride. Since silver chloride is soluble in dilute ammonia, the cloudiness disappears. 4. What is the difference between a sulfate ion and a sulfite ion? How can these two ions be distinguished from each other in the laboratory? The sulfate ion has the formula SO4 2- and the formula for the sulfite ion is SO32-. The two ions are distinguished by adding a few drops of barium chloride solution to a solution of each ion. A white precipitate is observed in each case. Add dilute hydrochloric acid to the precipitate. The cloudiness disappears in the case of the sulfite ion but remains in the case of the sulfate ion. 5. Why is it better to invent a fresh lime water solution before using it to confirm the presence of carbon dioxide? It is better to invent a fresh solution for lime water reacts with carbon dioxide in the air, and this makes the solution cloudy. 6. In test for the presence of nitrate association, why should it be taken extreme care when using concentrated sulfuric acid? Concentrated sulfuric acid is very corrosive and causes the skin burns. 4Leaving certificate chemistry: student laboratory notebook (overwhelming responses) Experiment 10.1. Measure the relative molecular mass of a volatile liquid. Questions about this experiment 1. Defining the term relative molecular mass? The relative molecular mass of a mixture is the mass of a molecule of that compound compared to a twelfth of the mass of the carbon isotope-12. 2. A volatile liquid is used in this experiment. What does the volatile term mean? A volatile liquid is one with a low boiling point (evaporating easily). 3. In addition to the liquid used in this experiment, give another example of a suitable liquid. Examples of suitable liquids are propane, cyclings, Estano and Ethanol. 4. What does the term s.t.p. represent? Standard temperature and pressure. The standard temperature is 273 k and the standard pressure is 1 105 Pascal. 5. Why should a very small hole be perforated in aluminum foil? The small hole is necessary to allow excess steam to escape the flask. 6. Which units are normally used to measure atmospheric pressure? Pascals (new per square meter) or PA symbols (N M-2). 7. What must be done to ensure that all the volatile liquid has completely vaporized? The flask must be left in boiling water for several minutes. 8. Indicate a source of error in this etnellob etrellob auca'llen attechsaif al eraiscal noN .otatseni ocinoc ocsaif led emulov li erarusM .otuicssa e otitup ais ksalf led onretni'li ehc eritnarag rep eracitnemid .erorre id itno? 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